

- 1 -

POROUS BODIES AND METHOD OF PRODUCTION THEREOF

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Technical Field

The present invention relates to water soluble or dispersible porous bodies and to methods of producing such porous bodies.

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Background to the Invention

Certain surfactants, including, for example, the >C8 primary  
15 alkyl sulphate ('PAS') surfactants having utility in detergent compositions, are difficult to dissolve in water at low temperatures. This may be due to the surfactant forming a viscous phase on contact with water and this phase can act as a mixing barrier, hindering further dissolution of surfactant. It may also  
20 be that the crystalline form of surfactant is very stable at low temperatures. While it would be advantageous, for environmental reasons, to be able to use PAS and certain other surfactants for laundering, and even more advantageous if this could be done at low temperatures, the dissolution kinetics of PAS have precluded  
25 its widespread use in low temperature washing.

Our co-pending international patent application PCT/GB03/03226 describes the formation of porous beads comprising a three dimensional open-cell lattice of a water-soluble polymeric  
30 material with an average bead diameter in the range 0.2 to 5mm. These are typically 'templated' materials formed by the removal

- 2 -

of a non-aqueous dispersed phase from a high internal phase emulsion. The beads are freeze-dried to remove the bulk of the aqueous phase. This leaves a 'skeletal' form of the emulsion behind. The beads dissolve rapidly in water and have the remarkable property that a water insoluble component dispersed in the emulsion prior to drying can also be dispersed in water on solution of the beads. Surfactant is present as an emulsifier.

#### 10 Brief Description of the Invention

We have now determined that effectively polymer-free systems, i.e. which comprise a surfactant but little or no polymer can provide highly porous bodies which disperse rapidly on contact with water, even at low temperatures.

In accordance with a first aspect of the invention, there is provided water dispersible or water soluble porous bodies comprising a three dimensional open-cell lattice containing:

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(a) less than 10% by weight of water-soluble polymeric material other than a surfactant, and

(b) 5 to 95% by weight of a surfactant, said porous bodies having an intrusion volume as measured by mercury porosimetry (as hereinafter described) of at least about 3 ml/g

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These porous bodies are particularly advantageous especially when they contain a relatively high level of surfactant as this promotes rapid dissolution of the surfactant and dispersion of

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- 3 -

any optional water insoluble material for which the surfactant is a carrier.

The present invention also provides a method for the  
5 preparation of said porous bodies which comprises the steps of:

(I) cooling an emulsion of:

10 a) An aqueous phase comprising the surfactant, and,

b) An immiscible second phase,

15 to a temperature at which the continuous phase becomes solid, and

(II) subsequently removing the bulk of the continuous and dispersed phases.

20 The cooled emulsion retains its structure when the bulk of the phases are removed (preferably by freeze drying) leaving a solid, surfactant-containing lattice. This lattice retains its structure provided that the ambient temperature does not rise above its melting point. The lattice so produced is  
25 characterised by a large surface area, which greatly assists the solution of its components. This improved solution rate is particularly beneficial when the surfactant is being used for delicate cleaning tasks such as for cleaning delicate fabrics or where only cold water is available for use in the  
30 cleaning process.

### Detailed Description of the Invention

In order that the present invention may be better understood and carried forth into practice, it is described below with reference  
5 to various preferred features and particular embodiments.

#### **Surfactant:**

It is preferred that a substantial part of the structuring of the  
10 bodies, i.e. the lattice, is provided by the surfactant.  
Consequently, the surfactant is preferably a solid per-se at temperatures encountered during product storage, i.e. at temperature below 30 Celsius, preferably at temperatures below 40 Celcius. In the alternative, the surfactant may form a solid over  
15 an appropriate temperature range in the presence of other materials present in the composition, such as builder salts.

The surfactant may be non-ionic, anionic, cationic, or zwitterionic.

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Examples of suitable non-ionic surfactants include ethoxylated triglycerides; fatty alcohol ethoxylates; alkylphenol ethoxylates; fatty acid ethoxylates; fatty amide ethoxylates; fatty amine ethoxylates; sorbitan alkanoates;  
25 ethylated sorbitan alkanoates; alkyl ethoxylates; pluronics™; alkyl polyglucosides; stearyl ethoxylates; alkyl polyglycosides.

Examples of suitable anionic surfactants include alkylether  
30 sulfates; alkylether carboxylates; alkylbenzene sulfonates; alkylether phosphates; dialkyl sulfosuccinates; alkyl

- 5 -

sulfonates; soaps; alkyl sulfates; alkyl carboxylates; alkyl phosphates; paraffin sulfonates; secondary n-alkane sulfonates; alpha-olefin sulfonates; isethionate sulfonates.

5 Examples of suitable cationic surfactants include fatty amine salts; fatty diamine salts; quaternary ammonium compounds; phosphonium surfactants; sulfonium surfactants; sulfonxonium surfactants.

10 Examples of suitable zwitterionic surfactants include N-alkyl derivatives of amino acids (such as glycine, betaine, aminopropionic acid); imidazoline surfactants; amine oxides; amidobetaines. Mixtures of surfactants may be used.

15 The preferred surfactants are primary and secondary alcohol sulphates containing greater than C8 chain length, more preferably, the materials known commercially as 'cocoPAS'.

**Product Form:**

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The bulk density of the porous bodies is preferably in the range of from about 0.01 to about 0.2 g/cm<sup>3</sup>, more preferably from about 0.02 to about 0.09 g/cm<sup>3</sup>, and most preferably from about 0.03 to about 0.08 g/cm<sup>3</sup>.

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The intrusion volume of the porous bodies as measured by mercury porosimetry (as hereinafter described) is at least about 3 ml/g, more preferably at least about 4 ml/g, even more preferably at least about 5ml/g, and most preferably at

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least about 6 ml/g. For example, the intrusion volume may be from about 3 ml/g to about 30 ml/g, preferably from about

- 6 -

4 ml/g to about 25ml/g, more preferably from about 10 ml/g to about 20ml/g.

Intrusion volume provides a very good measure (in materials  
5 of this general type) of the total pore volume within the porous bodies of the present invention.

The porous bodies may be in the form of powders, beads or moulded bodies. Powders may be prepared by the  
10 disintegration of porous bodies in the form of beads or disintegration of bodies during other stages of the production process.

**Porous bodies as carriers:**

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The porous bodies of the present invention optionally include within the lattice hydrophobic materials to be dispersed when the bodies are dispersed in an aqueous medium. Dispersion into an aqueous medium of such hydrophobic materials is much improved.

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The hydrophobic materials may be incorporated into the lattice by dissolving them in the discontinuous oil phase of an oil-in-water emulsion from which the lattice is made.

25 The present invention also includes, in a further aspect, solutions or dispersions comprising surfactant and a hydrophobic material formed by exposing to an aqueous medium porous bodies according to the present invention, wherein said bodies comprise the hydrophobic material.

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- 7 -

There are many instances in personal care products such as deodorants, skin and hair cleaning or care products or in household products such as laundry cleaning and care products or household cleaning or care products for hard and soft surfaces  
5 where it is desirable to administer hydrophobic materials in an aqueous environment. Because of the hydrophobic nature of these materials they are often reluctant to disperse in an aqueous environment. The use of the porous bodies of the present invention facilitates this dispersion and in many cases enables  
10 hydrophobic materials to be dispersed more effectively than previously.

It may be required to disperse the hydrophobic materials at the point where the product is being used. In this case the porous  
15 bodies of the present invention will be contained in the product until it is used by exposing it to an aqueous environment, at which time the water-soluble/dispersible lattice of the porous body will break down releasing the hydrophobic material.

20 The porous bodies of the present invention may be used to introduce hydrophobic materials into products, for example, liquid products during the manufacture of the products. In this case the lattice of the porous bodies of the present invention will break down when the porous bodies contact an  
25 aqueous environment during manufacture releasing the hydrophobic material in a form in which it can be more readily incorporated into the product being manufactured.

The porous bodies of the present invention may be used to  
30 transport materials to sites where they can be incorporated into products. By converting liquid products into porous



- 8 -

bodies the need to transport large amounts of liquids can be avoided resulting in significant cost savings and safer transport of materials which are potentially hazardous when transported in a liquid form. Materials which would be potentially unstable if stored or transported in liquid form may be incorporated into the porous bodies of the present invention and stored or transported with less risk of degradation.

10 The incorporation of potentially unstable hydrophobic materials, for example vaccines, vitamins or perfumes, into the porous bodies of the present invention may protect them from degradation during storage prior to use.

15 Some specific examples of products in which the porous bodies of the present invention may be used are given below. These are given as examples only and are not intended to limit the applicability of the present invention. Those skilled in the art will however realise that the porous  
20 bodies of the present invention will have utility in other areas not specifically exemplified herein.

Hydrophobic materials that are released from the porous bodies of the present invention at the time of use may  
25 include:-

- antimicrobial agents, for example: triclosan, climbazole, octapyrox, ketoconazole, phthalimoperoxyhexanoic acid (PAP), quaternary ammonium compounds, colloidal silver,  
30 zinc oxide.



- 9 -

- antidandruff agent for example: zinc pyrithione
- skin lightening agents for example 4-ethylresorcinol
- 5 • fluorescing agents for example: 2,5-bis(2-benzoxazolyl) thiophene for use on fabrics (such as cotton, nylon, polycotton or polyester) in laundry products
- skin conditioning agents, for example cholesterol
- 10 • antifoaming agents for example isoparaffin
- hair conditioning agents for example quaternary ammonium compounds, protein hydrolysates, peptides, ceramides and
- 15 hydrophobic conditioning oils for example hydrocarbon oils such as paraffin oils and/or mineral oils, fatty esters such as mono-, di-, and triglycerides, silicone oils such as polydimethylsiloxanes (e.g. dimethicone) and mixtures thereof
- 20 • fabric conditioning agents for example quaternary ammonium compounds having 1 to 3, preferably 2 optionally substituted (C8-C24) alk(en)yl chains attached to the nitrogen atom by one or more ester groups; hydrophobic
- 25 monoparticles such as a sucrose polyester for example sucrose tetra-tallowate; silicones for example polydimethylsiloxane
- thickening agents for example hydrophobically modified
- 30 cellulose ethers such as modified hydroxyethylcelluloses

- 10 -

- dyes for example dyes intended to change the colour of fabrics, fibres, skin or hair.
- UV protecting agents such as sunscreens for example octyl methoxycinnamate (Parsol MCX), butyl methoxydibenzoylmethane (Parsol 1789) and benzophenone-3 (Uvinul M-40), ferulic acid.
- bleach or bleach precursors for example 6-N-phthalimidoperoxyhexanoic acid (PAP) or photobleaching compounds. Dispersing the bleach from the porous bodies of the present invention results in the bleach being more finely dispersed and reduces the spot damage seen when larger particles of the bleach contact a fabric
- antioxidants for example hydrophobic vitamins such as vitamin E, retinol, antioxidants based on hydroxytoluene such as Irganox or commercially available antioxidants such as the Trollox series.
- insecticides, pesticides, herbicides that are stored as solid compositions before use but which are made up into liquid for spraying onto animals or crops
- perfumes or flavourings or precursors thereto
- pharmaceutically or veterinary active materials. There is a need for pharmaceutical compositions which can be taken by the consumer without the need to ingest the composition with a drink such as water. These compositions interact

- 11 -

with the moisture in the oral cavity to release the active ingredient which is then ingested by the consumer. By incorporating the pharmaceutically or veterinary active molecule in the porous bodies of the present invention, pharmaceutical compositions which meet this need can be prepared.

- In a similar way to that described above pharmaceutical and veterinary active ingredients may be formulated so that they release the active material into the nasal, ocular, pulmonary or rectal cavities or on the skin where they may act topically or they may be absorbed transdermally to act systemically
- By using the appropriate polymeric material in the lattice of the porous bodies of the present invention, porous bodies can be made that remain intact until the conditions (for example temperature or pH) change to those under which dispersion can occur. Thus dispersion can be delayed until a certain temperature has been reached or until the pH has changed to a suitable value such as would occur as the porous bodies pass down the GI tract. The acidity in the GI tract reduces down the GI tract and porous bodies which disperse hydrophobic actives only when the porous bodies are exposed to higher pH conditions enable pharmaceutically or veterinary active materials to be released only in the intestine having passed through the stomach intact.

Examples of situations where the porous bodies of the present invention are used to incorporate a hydrophobic

- 12 -

material into a product during the manufacture of that product include:-

- 5     • the introduction of hydrophobic materials such as fluorescers; enzymes; bleaches; hydrophobic polymers for example hydrophobically modified polyacrylates, silicones, hydrophobically modified polyvinylpyrrolidone, sulpha alkyl polysaccharides, Jaguar and JR polymers; fatty alcohols or acids; dyes for example shading dyes or black  
10     dyes for colour recovery into laundry products.
- 15     • the use of porous bodies according to the present invention containing hydrophobic dyes in the manufacture of water soluble inkjet compositions.
- 20     • the introduction of porous bodies containing different hydrophobic materials enables a manufacturer to produce a single base formulation into which the desired hydrophobic materials may be introduced by the use of the appropriate porous body of the present invention.
- 25     • the use of porous bodies containing hydrophobic polymers which disperse into water as the lattice breaks down to form a latex. The use of such latexes containing appropriate hydrophobic polymers deposited onto fabric  
30     imparts crease resistance or easy-iron properties to the fabric.

The porous bodies of the present invention may include within the  
30     lattice, water soluble materials which will be dispersed when the

- 13 -

bodies are dispersed in an aqueous medium. The water soluble materials may be incorporated into the lattice by dissolving them in the liquid medium from which they are made.

5 Examples of suitable water soluble materials include:-

- Water soluble vitamins such as vitamin C;
- water soluble fluorescers such as the 4,4'-bis(sulfo-  
10 styryl)biphenyl disodium salt (sold under the trade name Tinopal CBS-X;
- activated aluminium chlorohydrate;
- 15 • transition metal complexes used as bleaching catalysts;
- water soluble polymers such as polyesters isophthalic acid), gerol, xanthan gum, or polyacrylates; diethylenetriaminepentaacetic acid (DTPA);

20

or mixtures thereof

The porous bodies of the present invention may include within the lattice, materials which will be dispersed as  
25 very small particles when the polymeric bodies are dispersed in an aqueous medium. These materials may be incorporated into the lattice by dissolving or dispersing them in the liquid medium from which the porous bodies are made. If the particles are less than 1 micron, preferably less than 0.5  
30 micron and they are incorporated into skincare products then

- 14 -

the particles will not be felt by the user as the dispersed porous bodies are applied to the skin.

**Water soluble polymers:**

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Preferably the porous bodies of the present invention contain less than 5%wt more preferably less than 3% by weight of water-soluble polymeric material.

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When present, the polymeric material is a material that would be considered as "water soluble" by those skilled in the art i.e. if it forms a homogeneous solution in water. Water soluble polymers generally possess pendant polar or ionizable groups (e.g.  $-C=O$ ,  $-OH$ ,  $-N(R_1)(R_2)$  in which  $R_1$  and  $R_2$ , which may be the same or different, are independently H or (C1 to C4)alkyl,  $-N(R_3)(R_4)(R_5)^+$  in which  $R_3$ ,  $R_4$  and  $R_5$  which may be the same or different, are independently H or (C1 to C4)alkyl,  $-CON(R_6)(R_7)$  in which  $R_6$  and  $R_7$ , which may be the same or different, are H or (C1 to C4) alkyl, -  
15  $CH_2CH_2O-$ ,  $-CO_2H$  or salts thereof,  $-SO_3H$  or salts thereof groups) on a backbone chain which may be hydrophobic.  
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If water-soluble polymeric materials are incorporated into the porous bodies of the present invention, the time taken for the bodies to dissolve or disperse may be significantly reduced. For detergent applications, the nature of the lattice should be such that the dissolution or dispersion of the bodies preferably occurs in less than three minutes, more preferably less than two minutes, most preferably less  
30 than one minute.

- 15 -

Examples of water soluble polymeric materials include:-

- 5 (a) natural polymers (for example naturally occurring gums such as guar gum or locust bean gum or a polysaccharide such as dextran or cellulose;
- 10 (b) cellulose derivatives for example xanthan gum, xyloglucan, cellulose acetate, methylcellulose, methylethylcellulose, hydroxyethylcellulose, hydroxyethylmethylcellulose, hydroxy-propylcellulose, hydroxypropylmethylcellulose (HPMC), hydroxy-propylbutylcellulose, ethylhydroxyethylcellulose, carboxy-methylcellulose and its salts (eg the sodium salt - SCMC), or carboxymethylhydroxyethylcellulose and  
15 its salts (for example the sodium salt);
- (c) homopolymers of any one of the monomers listed in Table 1 below;
- 20 d) copolymers prepared from two or more monomers listed in Table 1 below;
- (e) mixtures thereof



- 16 -

Table 1

5	vinyl alcohol,
	acrylic acid,
	methacrylic acid
	acrylamide,
	methacrylamide
	acrylamide methylpropane sulphonates
10	aminoalkylacrylates
	aminoalkylmethacrylates
	hydroxyethylacrylate
	hydroxyethylmethacrylate
	vinyl pyrrolidone
15	vinyl imidazole
	vinyl amines
	vinyl pyridine
	ethyleneglycol
	ethylene oxide
20	ethyleneimine
	styrenesulphonates
	ethyleneglycolacrylates
	ethyleneglycol methacrylate

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When the polymeric material is a copolymer it may be a statistical copolymer (heretofore also known as a random copolymer), a block copolymer, a graft copolymer or a  
30 hyperbranched copolymer. Comonomers other than those listed in Table 1 may also be included in addition to those listed if their presence does not destroy the water soluble or

- 17 -

water dispersible nature of the resulting polymeric material.

Examples of suitable homopolymers include polyvinylalcohol,  
5 polyacrylic acid, polymethacrylic acid, polyacrylamides  
(such as poly-N-isopropylacrylamide), polymethacrylamide;  
polyacrylamines, polymethylacrylamines, (such as  
polydimethylaminoethyl-methacrylate and poly-N-  
morpholinoethylmethacrylate, polyvinyl-pyrrolidone,  
10 polyvinylimidazole, polyvinylpyridine, polyethylene-imine  
and ethoxylated derivatives thereof.

#### Method of Preparation:

15 As noted above, one method suitable for preparing the porous  
bodies comprises the steps of: cooling a surfactant-  
containing oil-and-water emulsion to a temperature at which  
the continuous phase becomes solid, and subsequently  
removing the bulk of the continuous and dispersed phases.

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Accordingly a further aspect of the present invention, there  
is provided a method the preparation of water dispersible or  
water soluble porous bodies comprising a three dimensional  
open-cell lattice containing:

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(a) less than 10% by weight of a water soluble polymeric  
material and

(b) 5 to 90% by weight of a surfactant,

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said porous bodies having an intrusion volume as measured by  
mercury porosimetry (as herein described) of at least about

- 18 -

3 ml/g with the proviso that said porous body is not a spherical bead having an average bead diameter of 0.2 to 5mm comprising the steps of:

- 5 a) providing an intimate mixture of the surfactant in a liquid medium
- b) providing a fluid freezing medium at a temperature effective for rapidly freezing the liquid medium;
- 10 c) cooling the liquid medium with the fluid freezing medium at a temperature below the freezing point of the liquid medium for a period effective to rapidly freeze the liquid medium; and
- 15 (d) freeze-drying the frozen liquid medium to form the bodies by removal of the liquid medium by sublimation.

The intimate mixture of the surfactant in the liquid medium may  
20 be an oil-in-water emulsion comprising a continuous aqueous phase containing the polymeric material, a discontinuous oil phase and the surfactant.

When the porous body is to be in the form of a powder the  
25 cooling of the liquid medium may be accomplished by spraying the liquid medium in atomised form into the fluid freezing medium.

When the porous body is to be in the form of beads the  
30 cooling of the liquid medium may be accomplished by dropping drops of the liquid medium into the fluid freezing medium.

- 19 -

Porous bodies in the form of moulded bodies may be made by pouring the liquid medium into a mould and cooling the liquid medium by the fluid freezing medium. In a preferred process of the invention to make moulded bodies, the liquid  
5 medium is poured into a pre-cooled mould surrounded by fluid freezing medium.

The frozen liquid medium may be freeze-dried by exposing the frozen liquid medium to high vacuum. The conditions to be  
10 used will be well known to those skilled in the art and the vacuum to be applied and the time taken should be such that all the frozen liquid medium present has been removed by sublimation.

15 In the case of moulded porous bodies freeze-drying may take place with the frozen liquid medium still in the mould. Alternatively, the frozen liquid medium may be removed from the mould and subsequently freeze-dried.

20 The freeze-drying step may be performed for up to around 72 hours in order to obtain the porous bodies of the present invention.

The above process preferably uses an oil-in-water emulsion  
25 comprising a continuous aqueous phase and a discontinuous oil phase. The surfactant usefully acts as an emulsifier for the emulsion.

Surfactants suitable for use as emulsifiers in oil-in-water  
30 emulsions preferably have an HLB value in the range 8 to 18. It is preferred that the surfactant is present in the liquid

- 20 -

medium in a concentration of about 1% to about 60% by weight. More preferably, the surfactant is present in the liquid medium in a concentration of about 2 % to about 40 % by weight and a yet more preferred concentration is about 5%  
5 to about 25% by weight.

The discontinuous oil phase of the oil-in-water emulsion preferably comprises a material which is immiscible with the continuous phase, which freezes at a temperature above the  
10 temperature which is effective for rapidly freezing the aqueous medium and which is removable by sublimation during the freeze drying stage.

The discontinuous oil phase of the emulsion may be selected  
15 from one or more from the following group of organic solvents:

- alkanes, such as heptane, n-hexane, isooctane, dodecane, decane;  
20
- cyclic hydrocarbons, such as toluene, xylene, cyclohexane;
- halogenated alkanes ,such as dichloromethane, dichloroethane, trichloromethane (chloroform),  
25 fluorotrichloromethane and tetrachloroethane;
- esters such as ethyl acetate;
- 30 • ketones such as 2-butanone;

- 21 -

- ethers such as diethyl ether;
- volatile cyclic silicones such as cyclomethicone;

5 and mixtures thereof

Preferably, the organic solvent comprises from about 10 % to about 95 % v/v of the emulsion, more preferably from about 20 % to about 60 % v/v. A preferred solvent is cyclohexane as  
10 the freezing point of cyclohexane is higher than that of water and the specific heat capacity for cyclohexane is much lower than that of water. This induces rapid freezing of the emulsion.

Preferably, the fluid medium is at a temperature below the  
15 freezing point of all of the components and is preferably at a much lower temperature to facilitate rapid freezing. The fluid freezing medium is preferably a liquified substance which is a gas or vapour at standard temperature and pressure. The liquified fluid freezing medium may be at its  
20 boiling point during the freezing of the liquid medium or it may be cooled to below its boiling point by external cooling means. The fluid freezing medium may be selected from one or more of the following group; liquid air, liquid nitrogen (b.p.  $-196^{\circ}\text{C}$ ), liquid ammonia (b.p.  $-33^{\circ}\text{C}$ ), liquified noble  
25 gas such as argon, liquefied halogenated hydrocarbon such as trichloroethylene, chlorofluorocarbons such as Freon (RTM), hexane, dimethylbutene, isoheptane or cumene. Mixtures of organic liquids and solid carbon dioxide may also be used as the fluid freezing medium. Examples of suitable mixtures  
30 include chloroform or acetone and solid carbon dioxide ( $-77^{\circ}\text{C}$  and diethyl ether and solid carbon dioxide ( $-100^{\circ}\text{C}$ ).

- 22 -

The fluid medium is removed during freeze drying, preferably under vacuum and is preferably captured for reuse. Due to the very low boiling temperature, inertness, ease of expulsion and economy, liquid nitrogen is the preferred  
5 fluid freezing medium.

The emulsions are typically prepared under conditions which are well known to those skilled in the art, for example, by using a magnetic stirring bar, a homogenizer, or a rotator  
10 mechanical stirrer.

The porous polymeric bodies produced usually comprise of two types of pores. One is from the sublimation of solid ice. The other kind of pore structure results from the  
15 sublimation of the oil phase.

The method for producing porous bodies according to the present invention, will now be more particularly described, by way of example only, with reference to the accompanying  
20 Examples.

### Examples

25 In the Examples that follow the dissolution time is measured as using a weighed sample of the polymeric bodies, which was stirred gently with water until the stirred mixture was clear to the eye. The time at which the mixture became clear to the eye was recorded as the dissolution time.

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- 23 -

The freeze-drier used was an Edwards Supermodulyo. This was operated with an average vacuum of 0.2mbar and at -50 °C.

5 Example 1 - Hydrophilic Surfactant

1g of sodium dodecyl sulfate (SDS) was dissolved in 5ml of water. To this aqueous solution was added 0.5mg of oil red in 5ml cyclohexane with vigorous stirring (using a RW11  
10 Basic IKA paddle stirrer). The emulsion formed was sprayed into liquid nitrogen using a trigger spray and the resulting frozen powder was freeze-dried to form a powder. The powder was highly porous, rapidly dissolving and dispersed the hydrophobic dye quickly into water to form a clear red  
15 'solution'.

Example 2 - Hydrophobic surfactant

20 An experiment was conducted in order to produce highly porous, rapidly dissolving oil-soluble powders that disperse a hydrophilic dye quickly into oil.

0.2g of sodium dioctyl sulfosuccinate was dissolved in 4 ml  
25 of toluene. To this organic solution was added 0.01g of direct yellow 50 in 4ml of water with vigorous stirring (as described in the previous example). The emulsion formed was sprayed into liquid nitrogen using a trigger spray and the resulting frozen powder was freeze-dried to form a powder.  
30 This powder dissolved readily into acetone to form a clear 'yellow' solution.

- 24 -

Example 3 - Dissolution of cocoPAS

A 12.5%wt aqueous solution of sodium lauryl sulfate (cocoPAS, EMAL 10PHD, ex Kao) was prepared by adding cocoPAS to deionised water. A sample of the solution (4 ml) was stirred with a type RW11 Basic IKA paddle stirrer and cyclohexane (6 ml) was added to form an emulsion having 75% v/v of discontinuous phase.

The beaker containing the emulsion was frozen in liquid nitrogen.

10

Once frozen, the emulsion was placed in a freeze-drier overnight. The same solution was made into beads by dripping the emulsion into liquid nitrogen, or made into a powder by spraying the emulsion into liquid nitrogen. Once frozen, the emulsion was placed in a freeze-drier overnight as before.

15

10mg of beads (or monolith, or powder) were added to 2 ml water, and stirred at 250 rpm. Using a stopwatch, the time to form a clear solution was determined. Results are given in the table below:

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Table 2: Dissolution of CocoPAS

Material	Dissolution at 5 °C	Dissolution at 20 °C
Na cocoPAS as received	>2min - not fully dissolved	1 min
Monolith	2 min 10 sec	30 sec
Beads	2 min	20 sec
Powder	35 sec	30 sec

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- 25 -

It can be seen that the dissolution time of the cocoPAS was significantly improved by forming it into the porous bodies of the present invention.

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Example 4 - CocoPAS plus fluorescer

9g of CocoPAS (sodium lauryl sulphate, EMAL 10PHD, ex Kao) was dissolved in 120 ml of water. To this aqueous solution  
10 was added 1.0g of Tinopal™ SOP in 120ml of dichloromethane with vigorous stirring. The emulsion formed was sprayed into liquid nitrogen using a trigger spray and the resulting frozen powder was freeze-dried to form a powder.

15 A sample was also made containing 30%wt loading of the fluorescer, i.e. 1g Tinopal and 2.3g of cocoPAS with 40ml water and 40ml dichloromethane.

Both samples dissolved quickly into water to give clear  
20 'solutions'.

Example 5 - CocoPAS plus fluorescer in products:

25 Samples of the 10% and 30% loaded bodies, prepared as disclosed in Example 4 were added to Brilhante (ex Gessy Lever, 2004) base formulation, without any fluorescer present. This was done in proportions to give final products with 0.05, 0.1 and 0.2%wt in the final formulation.  
30 The change in Ganz whiteness (Delta G) on cloth monitors, before and after washing, was compared with that obtained

- 26 -

with two control formulations:

a) Brilhante base with the addition of 0.1wt% fluorescer on product, added by dissolution into ethanol - the  
5 'ideal situation'.

b) Brilhante base with the addition of 0.1wt% fluorescer on product, dispersed in nonionic surfactant.

10 Measurements were performed for cotton, nylon and polyester.  
Results are shown in table 3 below:

Table 3

Ex.	Loading	%wt on prod.	carrier	Delta G Cotton	Delta G Nylon	Delta G Polyester
Con. 1	-	0.1	Ethanol	21.94	38.15	8.88
Con. 2	-	0.1	Nonionic	2.95	8.89	1.74
5a	10%	0.05	CocoPAS	18.34	26.08	6.90
5b	10%	0.1	CocoPAS	21.68	33.35	9.11
5c	10%	0.2	CocoPAS	27.09	36.59	10.47
5d	30%	0,05	CocoPAS	12.86	22.27	6.83
5e	30%	0.1	CocoPAS	19.69	30.34	8.97
5f	30%	0.2	CocoPAS	24.99	36.02	10.46

15

From the table it can be seen that the method of the invention gives deposition of fluorescer which is comparable with dosage of fluorescer in ethanol.